Stoichiometry Formula Sheet:

Mass, Moles, & Molar Mass: $n \rightarrow moles$ $m \rightarrow mass\ (g)$ $M_w \rightarrow molar\ mass\ (g/mol)$	$$n = \frac{m}{M_w} \qquad m = n * M_w \qquad M_w = \frac{m}{n}$$ $$1\ mol = 6.02\ x\ 10^{23} \qquad 1\ dozen = 12$$
Percent Composition:	$$mass\ \% = \frac{mass\ of\ element}{Total\ mass}\ x\ 100\%$$
Percent Yield:	$$\%\ Yield = \frac{Actual\ Yield}{Theoretical\ Yield}\ x\ 100\%$$
Percent Error:	$$\%\ Error = \frac{|Measured\ Value - Actual\ Value|}{Actual\ Value}\ x\ 100\%$$
Concentration: (Molarity) $n \rightarrow moles\ of\ solute$ $V \rightarrow Volume\ of\ Solution\ (Liters)$ $M \rightarrow Molarity\ (mol/L)$	$$M = \frac{n}{V} \qquad n = MV \qquad V = \frac{n}{M}$$
The Dilution Equation:	$$M_1 V_1 = M_2 V_2$$